

Rules for drawing Lewis Structures (PLEASE USE A PENCIL!!!!!! PLEASE)

1) COUNT valence electrons. This number of electrons must appear in your structure.

Note: Add electrons for anions and subtract electrons for cations.

2) DRAW skeleton structure with all single bonds. Element written first is usually in the center because it is least electronegative and most willing to share electrons.

3) ADD lone electron pairs so that octet rule is followed for all atoms except H (which only gets 2).

4) COUNT electrons in structure – if this number matches #1, then you're done!

If this number is greater than #1, then follow steps 5-7 below.

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5) CREATE a double bond between two of the atoms.

6) REMOVE 2 lone pair so that octet rule is followed.

7) RECOUNT ELECTRONS – if this number matches #1 you're done.

If this number is still greater than #1, then repeat steps 5-7 with different lone electron pairs. You may create another double bond or even a triple bond if necessary.

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9) PLACE BRACKETS around any ion and include the charge outside the brackets.