

## Rules for finding bond and molecular polarities

### Bond polarities:

1. Look up the electronegativity value of each element in the bond.
2. Subtract the electronegativities (make sure that this number is positive)
3. If the number is less than .2, then the bond is non-polar
4. If the bond is between 1.8 and .2, then the bond is polar.

### Molecular polarities:

1. Find the bond polarities.
2. If all the bonds are non-polar, then the molecule is non-polar.
3. If the bonds are polar, consider the molecular geometry:
  - a. If the molecule is bent or Trigonal pyramidal, then the molecule is polar
  - b. If the molecule is linear, Trigonal planar, or tetrahedral, and all of the bonds have the same degree of polarity, then the molecule is non-polar.
  - c. If the molecule is linear, Trigonal planar, or tetrahedral, and the bonds have different degrees of polarity, then the molecule is polar.