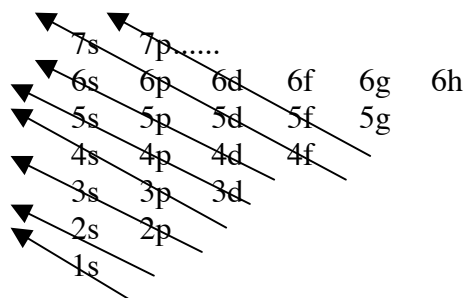


Electron Configurations

The Diagonal Rule:



electrons 2 6 10 14 18 22

- Write electron configurations for the following elements:
 - Sodium
 - arsenic
 - iron
 - gold
- Write the electron notation for iodine. What two sublevels are in the outer shell?
- How many electrons are in the highest energy level of the following?
 - antimony
 - tellurium
 - strontium
 - calcium
- Draw dot diagrams for:
 - oxygen
 - potassium
 - rubidium
 - indium
- How many d electrons are in :
 - vanadium
 - scandium
 - aluminum
 - arsenic
- How many p electrons are in:
 - aluminum
 - chlorine
 - neon
 - arsenic

Sublevel Sheet Two

- Write electron configurations for the following elements:
 - Potassium
 - Lithium
 - Cesium
 - Francium
- What do you notice about the outer shells for each of the above?
- How many electrons are in the highest energy level of the following?
 - antimony
 - arsenic
 - nitrogen
 - bismuth
- Draw dot diagrams for all of the elements in number 3. What do you notice about these diagrams?
- What conclusion could you make about the number of electrons in the outer shell of all of the elements in the same column on the periodic table?