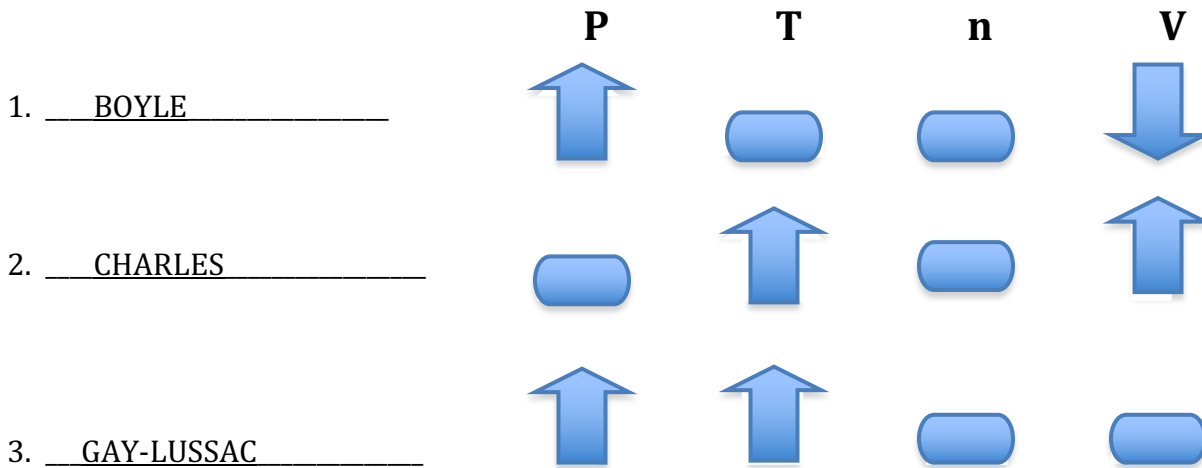


Gas Law Notes

Three Gas laws.



IF YOU LIKE FORMULAS:

$$T_1 V_2 P_2 = T_2 V_1 P_1$$

Units of pressure:

Atmospheres (atm) or Kilopascals (kPa) or millimeters of Hg (mm Hg)

Units of temperature: Kelvin (K) or Celcius (C)

STANDARD Pressures 1 atm 760 mm Hg (also called torr) 101.3 kPa

Standard Temperature 0 C 273 K

Ideal Gas Law Notes.....

$$P V = n R T$$

Temperature must be in KELVIN. K = C + 273

Volume must be in LITERS

Pressure can be in atm, or kPa, or mm Hg

R must match the units of pressure.

$$\mathbf{R \text{ Values}} = .08205 \frac{\text{l atm}}{\text{mole K}} \quad \text{or} \quad 8.31 \frac{\text{l kPa}}{\text{mole K}} \quad \text{or} \quad 62.4 \frac{\text{l mmHg}}{\text{mole K}}$$