

Periodic Table Trend Notes

Three trends + one

atomic radius --- atoms get smaller as you move left to right in a series/period
atoms get bigger as you move down a family/group

ionization energy. Smaller atoms hold their electrons tighter. SO... their ionization energy will be higher. Ionization energy is the energy required to remove an electron.

ionization energy increases as you move from left to right in a series/period
ionization energy decreases as you move down a family/group.

electronegativity Electronegativity is an atom's attraction for electrons. (technically, electrons in a bond.)

Small atoms attract electrons more UNLESS their outer shell is full.
SO..... electronegativity increases in a series/period
electronegativity decreases as you move down a family/group.

Activity. Big metals are more active. SO.....the most active metals are on the bottom left side of the periodic table

Small non metals are more active. SO.....the most active non-metals are smaller