

30 Equations for Fun and Points!!

Include State Symbols and simplify to NET IONIC REACTIONS where possible (single and double replacements). Assume constant oxidation states during synthesis reactions and for all decompositions except those of chlorates.

A couple of things to think about when considering state symbols:

- When not aqueous, ionic compounds should exist in what phase of matter?
- Potentially acidic compounds become acidic when placed in water
- Small molecular compounds will have weak IMF's, so they generally will exist as.....

A couple more things to consider when writing net reactions.

- Molecules do not dissociate unless they are strong acids
- Soluble ionic compounds dissociate in solution.
- Insoluble ionic compounds do not dissociate
- Dissociation only occurs in solution.
- When determining state of non-metal oxides, assume room temperature. You may need to look up melting and boiling points.

1. Aqueous aluminum chloride + aqueous silver nitrate

2. Solid nickel(III) carbonate is heated

3. Aluminum + aqueous zinc chloride

4. Chlorine + aqueous potassium iodide

5. Aqueous Sodium chloride + aqueous barium nitrate

